

# **INORGANIC MATERIALS SCIENCES** CEMENTS, PLASTERS, RAW MATERIALS

## Hydration of two calcium sulfates (plaster and anhydrite III)

### INTRODUCTION

• Four forms of calcium sulfate are known : dihydrate CaSO<sub>4</sub>·  $2H_2O$  (gypsum), hemihydrate CaSO<sub>4</sub>·  $\frac{1}{2}$  H<sub>2</sub>O (plaster) and two anhydrites CaSO<sub>4</sub> : natural and insoluble anhydrite and soluble anhydrite III (obtained by heating plaster to 200°C).

• Plaster and anhydrite III, which are soluble in water , are used in the building industry, because of their setting properties.

• Mixing calorimetry is a particularly interesting method for the study of short-term and long term heat effects during the setting of calcium sulfates.

#### **EXPERIMENT**

Samples : 1) Plaster (625 mg) + Water (500 mg) 2) Anhydride III (625 mg) + Water (500 mg) Crucible : Reversal mixing cell Heating mode : Isothermal 28°C



#### **RESULTS AND CONCLUSION**

At initial time, sample and water are separated by a lid in the mixing cell.

The mixing of the two compounds is carried out by reversing the calorimeter.

Two stages can be seen on the recorded thermograms :

. short term dissolution of the sample in water

. longer term setting of the mixture (hydration)

The rate of setting can be characterised by the time (to reach the top of the hydration peak.

Thus setting of plaster ( $t_{1/2}$  = 25 min ) is faster than the setting of anhydrite ( $t_{1/2}$  = 50 min).



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