

Measuring the oxygen stoichiometry (O/M ratio) of nuclear fuels with thermogravimetry

INTRODUCTION

Most nuclear fuels are oxides of radioactive metals like uranium or plutonium, or blends of these. The oxygen-tometal (O/M) ratio, or stoichiometry, is a critical parameter of nuclear fuel fabrication because it affects many of its properties. For instance, a too high O/M ratio may lead to a too low fuel's thermal conductivity, i.e. a less efficient heat transfer in the reactor.

Measuring the oxygen stoichiometry of nuclear fuels with thermogravimetry is a very accurate technique, based on measuring the subtle mass change from non-stoichiometric to

stoichiometric composition at a given temperature and under a controlled oxidative or reductive atmosphere.

EXPERIMENT

• Samples: pellets of mixed uranium-thorium oxide made using different protocols:

o (Th,3%U)O_{2+x} sintered under oxidizing and reducing atmosphere

- o (Th,3.75%U) $O_{2^{+x}}$ sintered under oxidizing and reducing atmosphere
- Instrument: Themys H2 TGA
- Temperature profile: heating from room temperature to 800°C

Under the final test conditions (Ar + 8%H₂ at 800 C), the dioxide is known to be stoichiometric, i.e. x=0. Using the mass loss recorded during the heating of the samples and applying the simple calculation shown on Figure 1, the O/M ratio of each sintered pellet could be

The O/M ratio of pellets sintered in air was slightly higher than that of pellets sintered in reducing atmosphere. The authors attributed this effect to the interstitial oxygen introduced into the fluorite lattice of

• Atmosphere: Ar + 8%H₂



RESULTS AND CONCLUSION

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Figure 1 – Idealized mass change signal for the O/M

Composition	O/M ratio	
	Oxidizing	Reducing
(Th,3%U)O₂	2.008	2.00
(Th,3.75%U)O₂	2.01	2.00

Table 1 – O/M ratio of the measured pellets

Reference : P.M. Khot et al. / Journal of Nuclear Materials 420 (2012) 1-8.

the mixed oxide structure during sintering in air.



measured.



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